

student exploration balancing chemical pdf

In a chemical reaction, reactants interact to form products. This process is summarized by a chemical equation. In the Balancing Chemical Equations Gizmo, look at the floating molecules below the initial reaction: $H_2 + O_2 \rightarrow H_2O$. 1. How many atoms are in a hydrogen molecule (H_2)? ____ 2. How many atoms are in an oxygen molecule (O_2)? ____ 3.

Student Exploration: Balancing Chemical Equations

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Balance: Turn on Show histograms. The equation is balanced when there are equal The equation is balanced when there are equal numbers of each type of atom represented on each side of the equation.

Student Exploration: Balancing Chemical Equations

Practice: Chemical reactions are generally classified into five groups, defined below. Balance each equation, using the Gizmo for help. Combination (or synthesis) – two or more elements combine to form a compound.

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Practice balancing chemical equations by changing the coefficients of reactants and products. As the equation is manipulated, the amount of each element is shown as individual atoms, histograms, or numerically.

Chemical Equations Gizmo : ExploreLearning

Chemical Equations. Practice balancing chemical equations by changing the coefficients of reactants and products. As the equation is manipulated, the amount of each element is shown as individual atoms, histograms, or numerically. Molar masses of reactants and products can also be calculated and balanced to

demonstrate conservation of mass.

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Balancing Chemical Equations "Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

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